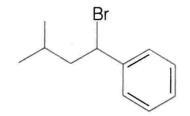
## MARKING SCHEME PRACTICE PAPER - 07

 (i) (b) In haloalkanes, C-atom with which halogen is attached is sp<sup>3</sup>-hybridised but in haloarene is sp<sup>2</sup>-hybridised.

Or

(b) Benzyl halogen compound is the one in which the halogen atom is bonded to an  $sp^3$ -hybridised carbon atom next to an aromatic ring, e.g.



(ii) (c) The IUPAC name of the compound is 4-bromopent-2-ene.

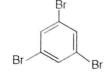
(1)

(1)

(1)

(1)

- (iii) (b) For dihalogen derivatives, the prefixes o, m, p are used in common system but in IUPAC system, the numerals 1, 2; 1, 3 and 1, 4 are used.
- (iv) (b)



Common name sym-tribromobenzene IUPAC name 1, 3,5-tribromobenzene

2. (i) (b) Amorphous solids soften over a range of temperature due to their short range order of arrangement of constituents. These can be moulded and blown into various shapes. Both assertion and reason are correct but reason is not the correct explanation of assertion.

Or

(d) Crystalline solids are anisotropic as these have different arrangement of particles along different directions. Hence, assertion is incorrect whereas reason is correct.

- (ii) (a) In covalent solids, atoms are held very strongly due to strong and directional nature of covalent bonds. Both assertion and reason are correct and reason is the correct explanation of assertion.
- (iii) (c) Although HCl and SO2 are polar molecular solids but they have neither free electrons nor free ions. Due to the absence of free electrons or ions these are non-conductors of electricity. Hence, assertion is correct and reason is incorrect.
- (iv) (b) Ionic solids have strong intermolecular forces of attraction which make them hard and brittle. Because of the presence of free ions, these can conduct electricity in molten state or in aqueous solution.

**3.** (c) Unit of rate constant = 
$$\frac{\text{time}^{-1}}{\text{concentration}^{(n-1)}}$$

where, n = order of reaction

Given, unit of rate constant = L mol<sup>-1</sup> s<sup>-1</sup>

$$\therefore \text{ L mol}^{-1} \text{ s}^{-1} = \frac{(\text{s})^{-1}}{(\text{L mol}^{-1})^{n-1}} = \frac{(\text{s})^{-1}}{(\text{L mol}^{-1})^{n-1}} = \text{s}^{-1} (\text{L mol}^{-1})^{n-1} \text{or} \qquad 1 = n - 1 \Rightarrow n = 2 \text{Order of reaction} = 2$$
(1)

.: Order of reaction = 2

4. (a) Formula of tetramminedichloroplatinum (IV) chloride is [Pt(NH<sub>3</sub>)<sub>4</sub>Cl<sub>2</sub>]Cl<sub>2</sub>. In aqueous solution, the above compound ionises as:

$$[Pt(NH_3)_4Cl_2]Cl_2 \xleftarrow{} [Pt(NH_3)_4Cl_2]^{2+} + 2Cl_2^{-}$$
  
So, the number of chloride ion is 2. (1)

5. (a) Catalyst used is  $V_2O_5$ 

2

(C)

$$SO_2(g) + O_2(g) \xrightarrow{V_2O_5} 2SO_3(g)$$

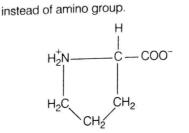
(b) On moving down the group H---X bond length increases so correct increasing order of their acidic strength is HF < HCl < HBr < HI (1)

6. (c) In the Hofmann bromamide degradation, amine is formed with side products Na<sub>2</sub>CO<sub>3</sub>, NaBr and H<sub>2</sub>O  $RCONH_2 + Br_2 + 4NaOH \longrightarrow RNH_2 + Na_2O_3$ 

$$+ 2NaBr + 2H_2O$$
 (1)

(c) 
$$(CH_3)_3N < (CH_3)_2NH < CH_3NH_2$$
, because the order of boiling points is 1° > 2° > 3°. (1)

7. (a) Proline has an imino (>NH) group



8. (a) Melting point of group 15 elements increases upto arsenic and then decreases upto bismuth.

So, the correct order is N < P < As > Sb > Bi

(d) N<sub>2</sub> molecule, N  $\equiv$  N, has one  $\sigma$ -bond and two (1)π-bonds.

- **9.** (b)  $CH_3CH = CH CH_3 + CHCl_3$ Chloroform But-2-ene C(OH)<sub>3</sub> NaOH hydrolysis CH-CH2CH3 COOH CH3-CH-CH2CH3 2- methylbutanoic acid
- 10. (d) Alkyl iodides are often prepared by the reaction of alkyl chlorides/bromides with Nal in dry acetone. This reaction is known as Finkelstein reaction.

 $RX + Nal \longrightarrow RI + NaX$ 

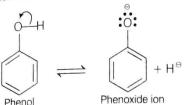
NaCl or NaBr thus formed is precipitated in dry acetone. It facilitates the forward reaction according to Le-Chatelier's principle.

Or

(c) The given reaction follows  $S_N^2$  mechanism and S<sub>N</sub>2 reactions are favoured in polar aprotic medium like DMSO, DMF... etc.

 $CH_3CH_2CH_2Br + NaCN \xrightarrow{DMF} CH_3CH_2CH_2CN + NaBr$ 

 (b) Phenols are much more acidic than alcohols due to stabilisation of phenoxide ion by resonance.



(1/2)

12. (c) Assertion is correct but Reason incorrect.

**Correct Reason** The total number of ions in solution increases on dilution but the number of ions per unit volume that carry current in solution decreases on dilution. (1)

(b) Both Assertion and Reason are correct statements but Reason is not the correct explanation of Assertion. (1)

Or

(a) Both Assertion and Reason are correct and Reason is the correct explanation of Assertion.

14. (c) Assertion is correct but Reason is incorrect.

**Correct Assertion** For  $d^4$  configuration, high spin complex with configuration  $t_{2g}^3 e_g^1$  is formed. (1)

- 15. (a) Both Assertion and Reason are correct statements and Reason is correct explanation of Assertion. (1)
- (b) Both Assertion and Reason are correct statements but Reason is not the correct explanation of Assertion.
   (1)
- 17. The depression in freezing point are in order:

 $CH_3COOH < CI_3CCOOH < F_3CCOOH$ Fluorine, being the most electronegative, has the highest negative inductive effect. As a result of this,  $F_3C$  COOH is the strongest acid while  $CH_3COOH$  is the weakest acid. Hence,  $CF_3$ —COOH ionises to the largest while  $CH_3COOH$  ionises to the minimum extent to give ions in their solutions in water. Greater the ions produced, greater is the depression in freezing point. Hence, the depression in freezing point is the maximum for the trifluroacetic acid and minimum for acetic acid. (2)

**18.** (i) As 
$$-\frac{d[B]}{dt} = 1 \times 10^{-2} \text{ mol L}^{-1} \text{ s}^{-1}$$
  
 $\therefore$  Rate  $= -\frac{1}{2} \frac{d[B]}{dt} = \frac{1}{2} \times 1 \times 10^{-2}$   
 $= 0.5 \times 10^{-2} \text{ mol L}^{-1} \text{ s}^{-1}$  (1)

(ii) Rate  $= -\frac{d[A]}{dt} = -\frac{1}{2}\frac{d[B]}{dt} = +\frac{1}{3}\frac{d[C]}{dt}$ Rate of change in concentration of A $= -\frac{d[A]}{dt} = -\frac{1}{2}\frac{d[B]}{dt} = 0.5 \times 10^{-2} \text{ mol } \text{L}^{-1} \text{ s}^{-1}$ Rate of change in concentration of C

$$=\frac{d[C]}{dt} \Rightarrow -\frac{3}{2}\frac{d[B]}{dt} = \frac{3}{2} \times 1 \times 10^{-2} = 1.5 \times 10^{-2}$$
(1)

**19.** The solution  $Sb_2S_3$  is a negatively charged solution. Greater the valency of the flocculating ion, greater is the coagulation. This is known as Hardy-Schulze rule. For a negatively charged solution,

$$AI^{3+} > Ca^{2+} > Na^+ \simeq NH_4^+$$

(2)

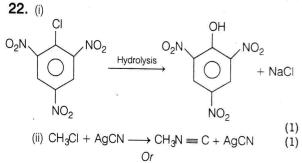
Hence, the most effective coagulating agent for  $Sb_2S_3$  solution is  $Al_2(SO_4)_3$ .

20. When SO<sub>2</sub> gas is passed through lime water first slowly, milky appearance arises due to formation of insoluble CaSO<sub>4</sub>. On passing excess SO<sub>2</sub> through this solution, milky appearance disappears because of formation of soluble Ca(HSO<sub>3</sub>)<sub>2</sub>.

$$Ca(OH)_{2} + SO_{2} \longrightarrow CaSO_{4} \downarrow + H_{2}O$$

$$CaSO_{3} + H_{2}O + SO_{2} \longrightarrow Ca(HSO_{3})_{2}$$
(2)

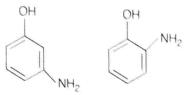
- 21. (i) The oxide of a transition metal in low oxidation state is basic, e.g. TiO, VO, MnO, FeO, Cu<sub>2</sub>O, NiO. The oxides of transition elements in higher oxidation states are acidic. e.g. V<sub>2</sub>O<sub>5</sub>, CrO<sub>3</sub>. (1)
  - (ii) This is because oxygen and fluorine are highly electronegative elements. (1)



Due to resonance in phenol, C—O bond of phenol has some partial double bond character. Partial double bond character strengthen the bond. So, it is difficult to break this C—O bond of phenol while the C—O bond of alcohol is purely single bond and comparatively weaker bond. So alkyl halides can be prepared by the reaction of alcohols with HCl in the presence of ZnCl<sub>2</sub> while aryl halides can not be prepared by reaction of phenol with HCl in the presence of ZnCl<sub>2</sub>. (1)

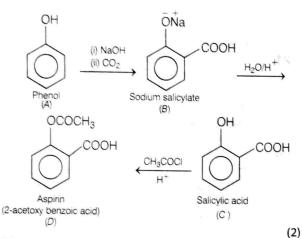
$$\begin{array}{c} C_{0}H_{5}OH \xrightarrow{HCI} No \text{ reaction} \\ Phenol & ZnCl_{2} \end{array} \xrightarrow{RCH_{2}OH} RCH_{2}OH \xrightarrow{HCI} RCH_{2}Cl + H_{2}O \\ Alcohol & ZnCl_{2} \xrightarrow{RCH_{2}Cl} Alkyl \text{ chloride} \end{array}$$
(1)

**23.** (i) The structures of *m*-aminophenol and *o*-amino phenol are as follows :



The  $--NH_2$  group is an electron releasing group. Thus, the presence of  $--NH_2$  group on the ring decreases the acid strength of phenol. The effect of any substituent is more significant when it is in *ortho* or *para* position than in *meta* position. Thus, *m*-amino phenol is stronger than o-aminophenol. (1)

(ii) In alcohols, the alkyl group is an electron releasing group. As a result, the electron density on oxygen increases. This effect along with the presence of two lone pairs of electrons on O-atom makes alcohols weakly basic. (1)



**24.** For NaCl, Z = 4, Molar mass,  $M = 58.5 \text{ g mol}^{-1}$ 

Density,  $d = 2.165 \, \text{g cm}^{-3}$ 

Density, 
$$d' = \frac{Z \times M}{a^3 \times N_A}$$
  
=  $\left(\frac{4 \times 58.5 \text{ g mol}^{-1}}{2.165 \text{ g cm}^{-3} \times 6.02 \times 10^{23} \text{ mol}^{-1}}\right)^{1/3}$   
 $a = (179.54)^{1/3} \times 10^{-8} \text{ cm} = 5.64 \times 10^{-8} \text{ cm}$  (3)

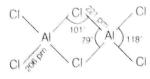
Distance between Na<sup>+</sup> and Cl<sup>-</sup> ions = a/2

$$=\frac{5.64 \times 10^{-8} \text{ cm}}{2}$$
$$= 2.82 \times 10^{-8} \text{ cm}$$

- (i) Due to inert-pair effect Bi(V) has a strong tendency to accept two electrons from any other species, thus acting as an oxidising agent. (1)
  - (ii) H<sub>3</sub>PO<sub>3</sub> has two oxygenated H-atoms. Since, P—OH hydrogen is ionisable, hence H<sub>3</sub>PO<sub>3</sub> is a diprotic acid.
     (1)

Or

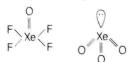
(i) The structural formula of Al2Cl6 is shown below ;



(ii) The structural formula of  $SF_4(g)$  is shown below. (1)



(1)
 (26. (i) The noble gas species which is isostructural with ICl<sub>4</sub> is XeF<sub>4</sub> and BrO<sub>3</sub> is isostructural with XeO<sub>3</sub>.
 Structures of XeF<sub>4</sub> and XeO<sub>3</sub> are as follows:



(ii) (a)  $\underset{\text{(Cold,dilute)}}{2\text{NaCH}} + \text{Cl}_2(g) \longrightarrow \text{NaCl}(aq) +$ 

2

30.

(1)

(b)  $XeF_6 + 3H_2O \longrightarrow XeO_3 + 6HF$  (1)

(b)  $XeF_6 + 3H_2O \longrightarrow XeO_3 + 6HF$  (1) **27.** (i) The outer electronic configuration of the given ions are:

V <sup>3+</sup>	Cu <sup>+</sup>	Sc <sup>3+</sup>
3d <sup>2</sup>	3ď <sup>10</sup>	3ď <sup>0</sup>
Coloured	Nod-d transition	
Fe <sup>3+</sup>	Co <sup>2+</sup>	MnO <sub>4</sub>
3d <sup>5</sup>	3d <sup>7</sup>	3d <sup>0</sup>
Coloured	Coloured	Coloured
	3d <sup>2</sup> Coloured Fe <sup>3+</sup> 3d <sup>5</sup>	3d <sup>2</sup> 3d <sup>10</sup> Coloured         Nod-d t           Fe <sup>3+</sup> Co <sup>2+</sup> 3d <sup>5</sup> 3d <sup>7</sup>

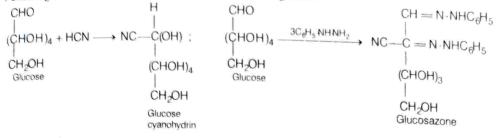
- (ii) (a) Due to lanthanoid contraction, the ionic size decreases on moving from La<sup>3+</sup> to Lu<sup>3+</sup>. As a result, the covalent character in bonding increases in the direction La<sup>3+</sup> → Lu<sup>3+</sup>. La<sub>2</sub>O<sub>3</sub> is more ionic and Lu<sub>2</sub>O<sub>3</sub> is more covalent.
  - (b) Radii of 4d and 5d-block elements are almost equal.
     (1)
- 28. (i) Amino acids may be acidic, basic or neutral depending upon the relative number of amino and carboxyl groups present in the molecule. Equal number of amino carboxyl groups makes it neutral, more number of amino than carboxyl groups makes it basic and more carboxylic groups as compared to amino groups makes it acidic. (11/2)
  - (ii) The amino acids exist as Zwitter ions
     (H<sub>3</sub>N—CHR—COO<sup>-</sup>) in aqueous solution. Due to this, dipolar salt like structure amino acids have strong dipole-dipole attractions.

Thus, their melting points are higher than corresponding halo acids which do not have salt like character. Secondly, they interact strongly with H2O due to their salt like character.

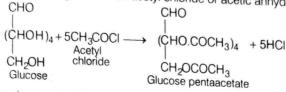
As a result, polar amino acids dissolve in polar solvents like H2O. Hence, the solubility of amino acids in water is higher than that of corresponding halo acids, which do not have any salt like character (11/2)

Or

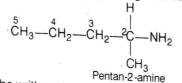
Following are the two reactions for the identification of glucose



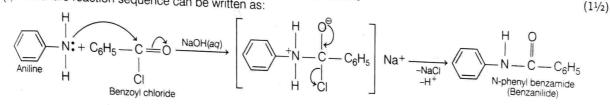
Reaction with acetyl chloride Reaction of glucose with acetyl chloride or acetic anhydride, gives glucose



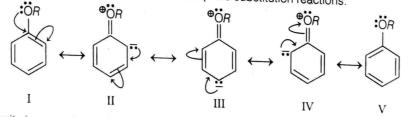
**29.** (i) HNO<sub>2</sub> reacts with  $C_5H_{13}N$  to give an alcohol, means the compound is primary amine.  $C_5H_{13}N$  means  $C_5H_{11}NH_2$ (11/2)(primary amine). Optically active alcohol means C<sub>5</sub>H<sub>11</sub>OH segment contain a chiral carbon.



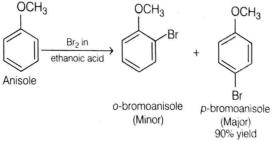
(ii) The entire reaction sequence can be written as:



**30.** (i) In aryl alkyl ethers, +*R*-effect of the alkoxy group (—OR) increases the electron density in the benzene ring,  $(1\frac{1}{2})$ thereby activating the benzene ring towards electrophilic substitution reactions.



(ii) Electron density is more at o-and p-positions so, o-and p-products are mainly formed during electrophilic  $(1\frac{1}{2})$ substitution reactions.



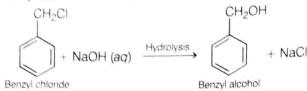
(11/2)

Or

(i) (a) Propene -----> Propan-2-ol

$$CH_{3} - CH = CH_{2} \xrightarrow{H^{*}/H_{2}O} CH_{3} - CH - CH_{3}$$
Propene
$$OH$$
Propan-2:ol
(1/2)

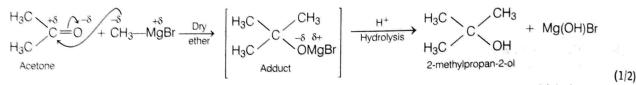
(b) Benzyl chloride ------ Benzyl alcohol



(c) Ethyl magnesium chloride ----> Propan-1-ol

$$\begin{array}{c} H \\ & & & \\ & & \\ H \\ & & \\$$

(d) Methyl magnesium bromide ------> 2-methylpropan-2-ol



(ii) The — OH group in alcohols is involved in intermolecular hydrogen bonding or H-bonding due to which they exist as associated molecules and hence, have higher boiling and melting points in comparison at hydrocarbons.

**31.** (i) 
$$2Cr(s) + 3Cd^{2+}(aq) \longrightarrow 3Cd(s) + 3Cr^{3+}(aq)$$

Thus,  $Cr(s)|Cr^{3+}(aq)|$  is anode and  $Cd^{2+}(aq)Cd(s)|$  is cathode.

Then, 
$$E_{cell}^{\circ} = E_{cathode}^{\circ} - E_{anode}^{\circ} = E_{(Cd^{2+}/Cd)}^{\circ} - E_{Cr/Cr^{3+}}^{\circ}$$

Taking the values of the standard electrode potentials,

$$E_{cell}^{\circ} = -0.40 \text{ V} - (0.74 \text{ V}) = +0.34 \text{ V}$$
  

$$\Delta_r G^{\circ} = -nF E_{cell}^{\circ} = -6 \times 96500 \times 0.34 \text{ J} \text{ mol}^{-1} - 196860 \text{ J} \text{ mol}^{-1}$$
  

$$\Delta_r G^{\circ} = -196.86 \text{ kg mol}^{-1}$$

The equilibrium constant of the cell reaction is obtained from equation,  $\Delta_r G^\circ = -2.303 RT \log K - 196860 \text{ J mol}^{-1} = -2.303 \times 8.314 \text{ J K}^{-1} \text{ mol}^{-1} \times 298 \text{ K} \times \log \text{ K}$ 

$$\log K = \frac{-196860 \text{ J mol}^{-1}}{-2.303 \times 8.314 \text{ JK}^{-1} \text{mol}^{-1} \times 298 \text{ K}} = 34.5 \implies K = 3.17 \times 10^{34}$$

(ii)  $Fe^{2+}(aq) + Ag^{+}(aq) \longrightarrow Fe^{3+}(aq) + Ag(s)$ 

Thus,  $Pt / Fe^{2+}$ ,  $Fe^{3+}$  electrode is anode and  $Ag^+ / Ag$  is cathode in this cell.

Then, 
$$E_{cell}^{\circ} = E_{cathode}^{\circ} - E_{anode}^{\circ} = E_{Ag^+/Ag}^{\circ} - E_{Fe^{3+}/Fe^{2+}}^{\circ}$$

Substituting E° values,

$$E_{cell}^{\circ} = 0.80 V - 0.77 V = 0.03 V$$
  

$$\Delta_r G^{\circ} = -nF E_{cell}^{\circ}$$
  

$$= -1 \times 96500 \times 0.03 \text{ J mol}^{-1} = -2895 \text{ J mol}^{-1} - 2.895 \text{ kJ mol}^{-1}$$

(21/2)

(1)

(1/2)

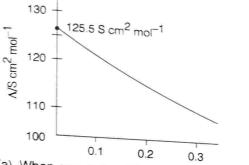
$$\log K = \frac{\Delta_{l} G_{cell}^{*}}{-2.303 \text{ RT}}$$
$$= \frac{-2895 \text{ J mol}^{-1}}{-2.303 \times 8.314 \text{ JK}^{-1} \text{mol}^{-1} \times 298 \text{ K}}$$
$$\log K = 0.507 \implies K = 3.2$$
(21/2)

Given data is as follows:

C (mol L <sup>-1</sup> )	0.001	0.01	0.02	0.05	
VC	0.032		0.02		
$10^2 \times \kappa (\mathrm{Sm}^{-1})$			0.14	0.22	0.32
	1.23/	11.85	23.15	55.53	106.74
$10^4 \times \kappa (\text{Sm}^{-1})$	1.237	11.85	23.15	55 53	106.74
Å	123 7	118 5	145 -	00.00	100.74
		110.0	115.7	111.1	106.74

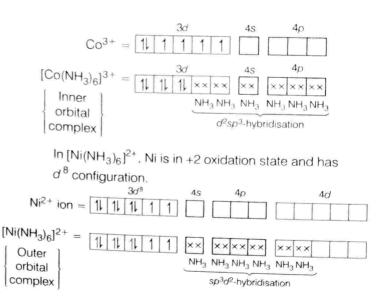
$$\hat{\Lambda} = \frac{1000 \times \kappa (\text{Sm}^{-1})}{C (\text{mol}/1)}$$

The  $\Lambda$ - $\sqrt{C}$  plot is shown in the following plot :



- 32. (i) (a) When one mole of CoCl<sub>3</sub>.6 NH<sub>3</sub> is mixed with AgNO<sub>3</sub>, three moles of AgCl are precipitated which indicates that three ionisable chloride ions in the complex are present. Hence, its structural formula is [Co(NH<sub>3</sub>)<sub>6</sub>]Cl<sub>3</sub>. (1)
  - (b) In both the complexes, Fe is in +2 state with the configuration  $3d^6$ , i.e. it has four unpaired electrons. As the ligands H<sub>2</sub>Oand  $CN^-$  possess different crystal field splitting energy ( $\Delta_0$ ) they absorb different components of visible light (VIBGYOR) for *d*-*d* transition. Hence, the transmitted colours are different.
  - (c) According to VBT,  $sp^{3}d^{2}$ -hybridisation is found in  $[Co[H_2O]^{2+}$  and its structure is octahedral. It is an outer orbital complex because in this complex two outer *d*-orbitals (*nd*) participate in hybridisation, and it has three unpaired electrons. It has a weak ligand (H<sub>2</sub>O). Its nature is paramagnetic. (1)

(ii) 
$$\ln [Co(NH_3)_6]^{3+}$$
, Co is present as Co<sup>3+</sup> and has 3d<sup>6</sup> configuration. In the presence of NH<sub>3</sub>, the 3d-electrons pair up leaving two d-orbitals empty to be involved in  $d^2sp^3$ -hybridisation forming inner orbital complex in the case of  $[Co(NH_3)_6]^{3+}$ .

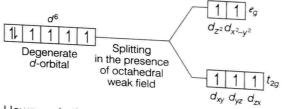


Since, (n - 1) d-orbitals are not available but the nd-orbitals are used in bond formation, i.e in hybridisation, the complex is called outer orbital complex.

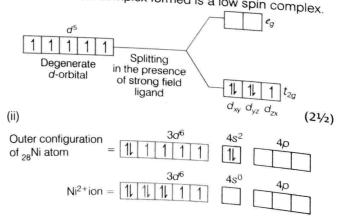
Or

(2)

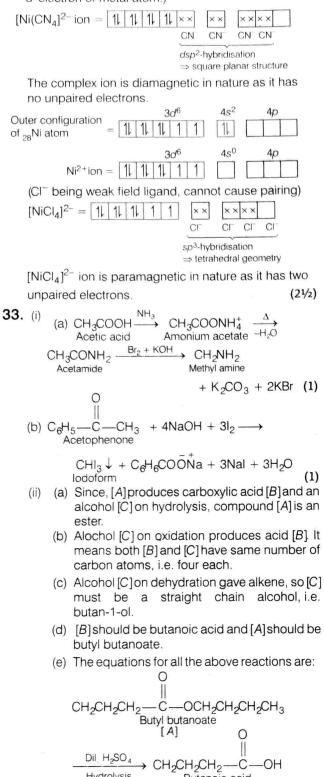
(i) Mn (II) ion has  $3d^{5}$  configuration. In the presence of H<sub>2</sub>O molecules (acting as weak field ligands), the distribution of these five electrons is  $t_{2g}^{3}e_{g}^{2}$ , i.e. all the electrons remain unpaired to form a high spin complex.



However, in the presence of  $CN^-$  (acting as strong field ligands), the distribution of these electrons is  $t_{2g}^5 e_g^0$ , i.e. two  $t_{2g}$  orbitals contains paired electrons while the third  $t_{2g}$  orbital contains one unpaired electron. The complex formed is a low spin complex.



(CN<sup>-</sup> being strong field ligand paired up the d-electron of metal atom.)

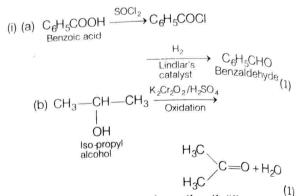


Hydrolysis

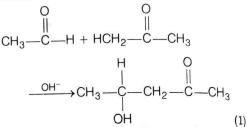
Butanoic acid [B]

> + CH3CH2CH2CH2OH Butan-1-ol [C]

 $CH_3CH_2CH_2CH_2OH \xrightarrow{[0]} CH_3CH_2CH_2COOH$ Butanoic acid (oxidation) Butan-1-ol [C]  $\xrightarrow{\text{Conc. H}_2\text{SO}_4} \text{CH}_3 \xrightarrow{\text{CH}_2} \text{CH}_2 \xrightarrow{\text{CH}_2} \text{CH}_2$ Dehydration (-H<sub>2</sub>O) But -1-ene Conc. H<sub>2</sub>SO<sub>4</sub> (3) Or



(a) Cross-aldol condensation If different (ii) carbonyl compounds having acidic hydrogen at a-position are condensed to give a. B-unsaturated compounds. Then, reaction is called cross-aldol condensaton.



(b) Cannizzaro's reaction Those aldehydes or ketones which do not have α-hydrogen atom reacts with conc. NaOH or KOH to form carboxylic acid salt and alcohol. This reaction is known as Cannizzaro's reaction.

e.g.

2HCHO + NaOH ----> HCOONa + CH₂OH Formaldehyde Sodium Methyl formate alcohol (1)

(c) HVZ (Hell-Volhard-Zelinsky reaction) a-hydrogen atom containing carboxylic acid in presence of red phosphorus reacts with chlorine or bromine to form a-halo substituted carboxylic acid. This reaction is known as HVZ reaction.

